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The Clausius-Clapeyron equation: $dP_i^{sat} = \frac{h_i^{vap}}{T^2} dT$ or $\ln P_i^{sat} = \frac{h_i^{vap}}{R} \left(\frac{1}{T} - \frac{1}{1373} \right) + \ln P_i^{sat}(1373)$ so $P_i^{sat} = 101 \text{ kPa} \exp \left(\frac{h_i^{vap}}{R} \left(\frac{1}{T} - \frac{1}{1373} \right) \right)$

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thermodynamic variables , i.e. $p=p(\rho,T)$
(6.1.1) where the functional relationship
in depends on the substance.

Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics
and Principles of Phase Equilibria In all
the preceding chapters we have focused

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primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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Chapter 6 Aqueous Thermodynamics. Overview. Understanding aqueous

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thermodynamics can be a daunting task. In this chapter we will describe some of the essential topics in aqueous thermodynamics and present them in a logical, relatively easy to understand manner.

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