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## **Covalent Bonding**

### **Key**

Covalent bond in which one atom contributes both bonding electrons; The shared electron pair comes from one of the bonding atoms (True or False) The negative charge of a polyatomic ion shows the number of electrons in addition to the valence electrons of the atoms present.

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**Chapter 8-Covalent  
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Chapter 9 Study Guide  
For Covalent Bond  
Answer - Chemistry  
Final Study Guide The  
covalent bond that  
arises from the overlap  
of two orbitals. (There  
is a greater Answer: ES  
2. 41 Chemical  
Bonding Test - Review  
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ANSWER KEY barium  
sulfite, ammonium

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phosphate. phosphorus

pentabromide.

magnesium sulfate.

calcium oxide.

phosphoric acid.

## **Chemistry Chapter 8 Covalent Bonding Study Guide Answer Key**

Ionic & Covalent

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1. An ionic bond is a bond between Metal (cation) & non-metal (anion). 2. What is the name given to the

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Key  
electrons in the highest occupied energy level of an atom? Valence electrons 3. How many electrons does nitrogen gain in order to achieve a noble-gas electron configuration? 3 4.

**Name: Ionic & Covalent Bonding - Study Guide An ionic bond ...**

covalent bond. in a covalent bond, the dissociation energy is

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Key released in the process of. exothermic reaction. When 2 or more atoms bond by means of electron sharing - the result is. a molecule. shared electrons are centered between the 2 atoms in covalent bonding, the attachment is called. a sigma bond.

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### Covalent Bonding

A double bond between two carbon atoms.

Notice that each carbon achieves eight electrons by this sharing. Because each shared pair constitutes a single covalent bond, the two shared pairs are called a double bond. The structure on the right side of Figure 8 shows this double bond of four shared electrons with two lines, and the left side of Figure

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Key Covalent Bonding  
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covalent bond is  
formed when two  
atoms share a pair of  
electrons. Covalent  
bonding occurs in most  
non-metal elements,  
and in compounds

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formed between non-metals. These shared electrons are found...

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1. Predict whether the following bonds will be ionic, polar covalent, or non-polar covalent. a.

C and O polar-

b. Na and Cl ionic c. B and Cl

d. H H non 2. Draw the

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Lewis Dot Diagram for  
the following ionic  
compounds: a. NaF b.  
CaCl<sub>2</sub> c. Al<sub>2</sub>O<sub>3</sub>

3. Differentiate between  
an ordinary covalent  
bond and a coordinate  
covalent bond. Give an

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A number in a chemical  
formula that tells the  
number of atoms in a  
molecule or the ratio of  
elements in a

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compound. covalent bond. A chemical bond that involves sharing a pair of electrons between atoms in a molecule. molecule.

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T182 Chemistry: Matter and Change Name

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Class CHAPTER 9

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CHAPTER 9 STUDY

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Naming Molecules In  
your textbook, read  
about how binary  
compounds and acids  
are named from their  
formulas.

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Pages 1 ...**

Welcome to this  
Bonding revision  
course. The course

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### Covalent Bonding

covers Structure and Bonding and how structure influences the properties and uses of substances which are assessed on Chemistry exams. The Course structure is a number of study notes and some visual aids to give you a comprehensive understanding of Bonding. The knowledge is re-enforced with a flashcard deck of key



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**Chapter 9 Covalent Bonding Study Guide Answer Key**

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Covalent bond in which one atom contributes both bonding electrons; The shared electron pair comes

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from one of the  
bonding atoms (True or  
False) The negative  
charge of a polyatomic  
ion shows the number

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and Chemistry Read  
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very strong bonds.

They're very important in biology because they're very stable and because most biological molecules are made with covalent bonds. These biological molecules...

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SHORT ANSWER  
Answer the following  
questions in the space  
provided. 1,



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**Chapter 6 Chemical  
Bonding Section 2  
Covalent Answer  
Key**

When atoms interact with each other, they can form molecules. Depending on how the electrons of these atoms interact, a covalent bond or ionic bond would be formed. Atoms bond in order to reach a more stable, lower energy state.

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